

Docket: 70-36 (M)  
Project: S-8

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Source & Special Nuclear Materials Br.

October 25, 1962

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UNITED NUCLEAR (MALLINCKRODT) LICENSE SM-33, SHIPPING CONTAINER  
FOR 20% ENRICHED  $UF_6$  (APPLICATION DATED 8/16/62) AND FOR  $\pm$  20%  
ENRICHED  $U_3O_8$  (APPLICATION DATED 10/15/62), DOCKET 70-36

United Nuclear requests extension of their license SM-33 to  
include a shipping container for 20% enriched  $UF_6$  and  $U_3O_8$   
enriched up to 20% having maximum tap densities of 1.5 grams/cc.

The container consists of a 6.12" I.D. 'Schedule 40' steel pipe  
centered in a standard 55-gallon drum birdcage formed by welding  
three such drums end to end. For a maximum of seven containers  
per shipment a total solid angle of 3.2 steradians will not be  
exceeded. We are in agreement with United Nuclear's analysis that  
the  $k_{eff}$  for a 6.12" diameter cylinder does not exceed 0.5 at an  
 $H/X$  of approximately 90 for the 20% enriched material, and consequently,  
a total solid angle of 3.2 steradians with the containers in a close  
packed hexagonal array, is safe. Commingling will be avoided through  
carrier certification.

In conclusion, we recommend approval of the applications, subject  
to Mr. Christian Beck's structural integrity evaluation. In this  
regard it is of course necessary that the evaluation criteria of  
10 CFR 70, as proposed, be invoked.

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Attachments:

1. Ltr fm UHC dtd 8/16/62
2. Ltr fm UHC dtd 10/15/62

cc - filed in "SNM Foreign Unenriched" - CFB

DLR:MS:CEB

CDLake/vj

10/26/62

B-108

Subject: United Nuclear Corp - SNM-23, Shipping Container for  
20% enriched  $UF_4$  (dtg. 8/16/62) and for 20% enriched  $U_3O_8$  (dtg.  
10/15/62). Docket 70-36

Containers: 6" diam. sched. 40 steel pipe (8'4" long) centered over  
a 3 (55 gal.) drum structure.

Array: 7 containers shipped in a resulting  $W = 3.2$   
(also - packed hexagonal array)

Calculations (Check:

Ref: DP-532 (table IV.16.)

$$3.77 \frac{\text{moles}}{L} \left( \frac{375g}{\text{mole}} \right) = \frac{1409g}{L} \text{ corresponds to } k_s = 1.84 \text{ for 20\% enriched } U_3O_8$$

$$\begin{aligned} 1 &= 237.5g \\ 2 &= 32g \\ 3 &= 34g \\ \hline 2 &= 17.5g \end{aligned}$$

$$1520g \frac{115g}{L} = \text{UF}_4 \text{ present}$$

$$P_{bulk} = 1.5g/cc$$

$$P_{theor} = 6.7g/cc$$

$$\frac{H}{U} = 18$$

$$\frac{H}{X} = 15 \times \frac{100}{20} = 90$$

would certainly  
be conservative  
for 20% enriched  $UF_4$

U. Nuc. assumed  $k_s = 2.05$

$$U_3O_8 = 8.39g/cc$$

$$\left( \frac{H}{X} \right)_{\text{wet}} = \frac{1 - \frac{1.5}{8.39}}{9} \left( \frac{235}{(1.5)(.847)(20)} \right) = \frac{(1.86)(235)}{(9)(.254)} = \frac{202}{228} = 89$$

Cellihan's Data: < 20% enriched  $U_3O_8$

$$1.9\% \rightarrow D_c(\text{unrefl.}) = 47cm @ \frac{1}{X} = 320$$

$$2\% \rightarrow D_c(\text{unrefl.}) = 54cm @ \frac{1}{X} = 450$$

United Nuclear  
Mallinbrodt  
Proj 5-8  
Doc 70-36

Cypl 8/16/62

Proposes shipment of 20% em  $UF_4$  having a bulk tap density of 1.5 g/cc

Assume  $\rho_0 = 10.9$  g/cc crystal density

For  $UF_4$   $\epsilon = 0.76$

$$\begin{aligned} (H/X)_{wet} &= \frac{1 - \frac{\rho}{\rho_0}}{9} \times \frac{235}{\rho \times \epsilon \times e} \\ &= \frac{1 - \frac{1.5}{10.9}}{9} \times \frac{235}{1.5 \times 0.76 \times 0.20} = 97.2 \end{aligned}$$

Assume  $\rho_0 = 6.7$  Crystal Density

$$(H/X)_{wet} = \frac{1 - \frac{1.5}{6.7}}{9} \times \frac{235}{1.5 \times 0.76 \times 0.20} = 89 \text{ or } U_3O_8$$

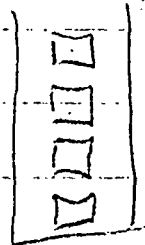
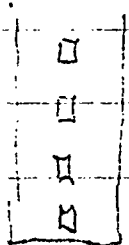
highly hyd  
dens.

low hyd  
dens.

Higher crystal density comp.  $\rho_0$

the more water, the

higher  $H/X$  for a  
given bulk  $\rho$



$$r = 15.4 \text{ cm.}$$

$$98\% \text{ } 245 \text{ cm}$$

$$Vol = 1.785 (15.4 \text{ cm})^2 (245 \text{ cm})$$

$$Vol = 46,350 \text{ cc}$$

$$\text{tap } \rho = 1.5 \text{ g/cc}$$

$$\therefore 69,525 \text{ g } UF_4$$

$$\text{Fuel} = 23\% \text{ enriched } UF_4$$

$$\text{tot kg } UF_4 = 69.5 \text{ kg}$$

$$\text{tot kg U} = 53 \text{ kg}$$

$$\text{" } U_{235} = 10.6 \text{ kg}$$

$$N_{25} = \frac{(10,600 \text{ g } U_{25}) \left( \frac{6.02 \times 10^{23} \text{ atoms}}{235 \text{ g } U_{25}} \right)}{46,350 \text{ cc}} = \frac{272}{46,350} = .00587 \times 10^{24} \text{ atoms/cc}$$

$$\frac{N_{15}}{N_{25}} = \frac{100-20}{20} = \frac{80}{20} = 4$$

? tot vol.  $H_2O$

$$\rho_{\text{crystal } UF_4} = 6.7 \text{ g/cc}$$

$$\left( \frac{H}{X} \right)_{\text{wet}} = \frac{1 - \frac{1.5}{6.7}}{9} \times \frac{235}{1.5 \times .76 \times .20} = 89$$

$$\therefore \frac{H_2O}{U_{235}} = 44.5$$

$$\text{g } H_2O = 43.5 \text{ g } U_{235}$$

$$= (43.5)(10,600)$$

$$= 470,000 \text{ g } H_2O = 470,000 \text{ cc } H_2O$$

$$\Sigma_{25} \text{ tot} = N_{25} \left[ \frac{637}{N_{25}} + \frac{N_{22}}{N_{25}} \frac{622}{N_{25}} + \frac{N_{H_2O}}{N_{25}} \frac{620}{N_{25}} \right]$$

$$= .00587 \times 10^{24} \text{ atoms/cc} [637 + 4(25) + 580(0.66)] \times 10^{-6} \text{ cm}^3$$

$$= (.00587)(1044.2)$$

$$= 0.615 \text{ cm}^{-1}$$

$$\frac{N_{H_2O}}{N_{25}} = \frac{470 \text{ kg } H_2O}{10.6 \text{ kg } U_{25}} = \frac{44.3}{1}$$

$$\frac{N_{22}}{N_{25}} = \frac{26.1}{0.45} = 580$$

$$\frac{637}{11.2} + \frac{622}{11.2} + \frac{580}{11.2} = 580$$

the  $\sigma_a$  of F is  $< 10 \text{ mb} \therefore$  neglect effect in  $\epsilon_a^{\text{tot}} \text{ calc.}$

$$\text{corrected } \epsilon_a^{\text{tot}} = \frac{.615}{1.13} = 0.545 \text{ cm}^{-1}$$

$$\text{Let } D = 0.162 \text{ cm.}$$

$$L^2 = \frac{D}{\epsilon_a^*} = \frac{0.162 \text{ cm}}{0.545 \text{ cm}^{-1}} = .3 \text{ cm}^2$$

$$f = \frac{\epsilon_{a,25}}{\epsilon_a^{\text{tot}}} = \frac{N_{25} \sigma_{25}}{\epsilon_a^{\text{tot}}} = \frac{(1.000587 \times 10^{24} \text{ atoms/cm}^3) (650 \times 10^{-24} \text{ cm}^2)}{0.615 \text{ cm}^{-1}}$$

$$f = \frac{.382}{.615} = .622$$

$$\text{Let } T = 33 \text{ cm}^2$$

$$\text{Buckling: } B_g^2 = \left( \frac{2.405}{7.2 + 3 \text{ cm}} \right)^2 = .054 \text{ cm}^{-2}$$

$$\eta = 2.05 ; \text{PF} = 1.0$$

$$k_{\text{eff}} = \frac{\eta f}{(1 + L^2 \epsilon_g^2) (1 + T B_g^2)} = \frac{(2.05) (.622)}{\left[ \frac{1 + (.3) (.054)}{1.0112} \right] \left[ \frac{1 + (33) (.054)}{2.8} \right]} = \frac{1.27}{2.8} = .454$$